

Chemistry – Chapter 19 Book problems #2: $[H^+]$ and $[OH^-]$ calculations

1. Classify each solution as acidic, basic, or neutral:

A. $[H^+] = 6.0 \times 10^{-10} \text{ M}$

B. $[OH^-] = 3.0 \times 10^{-2} \text{ M}$

C. $[H^+] = 2.0 \times 10^{-7} \text{ M}$

D. $[OH^-] = 1.0 \times 10^{-7} \text{ M}$

2. If the hydroxide-ion concentration of an aqueous solution is $1 \times 10^{-3} \text{ M}$, what is the $[H^+]$ in the solution? Is the solution acidic, basic, or neutral?