

Some of the problems below can be solved multiple ways. Use the ideal gas law to solve.

1. Calculate the volume (in L) of oxygen gas collected at STP if the sample has a mass of 2.67g.
2. Calculate the volume (in dm^3) of 101 grams of neon gas collected at STP.
3. 75.6 dm^3 of sulfur dioxide gas is collected at STP. What is its mass?
4. How many moles of gas will occupy a 562 cm^3 flask at -15.0°C and 88.7 kPa pressure?
5. What volume will be occupied by 0.766 mol of gas at 106 kPa and 15.5°C ?
6. A 759 cm^3 vessel contains 0.0925 mol of a gas at 98.6 kPa. What is its temperature (in K)?

