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10. Classify each solution as acidic, basic, or neutral:

A. $[H^+] = 6.0 \times 10^{-10}$ M

B. $[OH^-] = 3.0 \times 10^{-2}$ M

C. $[H^+] = 2.0 \times 10^{-7}$ M

D. $[OH^-] = 1.0 \times 10^{-7}$ M

11. If the hydroxide-ion concentration of an aqueous solution is 1×10^{-3} M, what is the $[H^+]$ in the solution? Is the solution acidic, basic, or neutral?