

Name _____

Empirical and Molecular Formula Worksheet

SHOW YOUR WORK ON A SEPARATE SHEET OF PAPER.

Write the empirical formula for the following compounds:

- 1) C_6H_6
- 2) C_8H_{18}
- 3) WO_2
- 4) $C_2H_6O_2$
- 5) $X_{39}Y_{13}$

- 6) A compound with an empirical formula of C_2OH_4 has a molar mass of 88 grams per mole. What is the molecular formula of this compound?
- 7) A compound with an empirical formula of C_4H_4O has a molar mass of 136 grams per mole. What is the molecular formula of this compound?
- 8) A compound with an empirical formula of $CFBrO$ has a molar mass of 253.8 grams per mole. What is the molecular formula of this compound?
- 9) A compound with an empirical formula of C_2H_8N has a molar mass of 46 grams per mole. What is the molecular formula of this compound?
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula of C_2H_4NO . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound?
- 11) Nitrogen and oxygen form an extensive series of oxides with the general formula of N_xO_y . One of them is a blue solid that comes apart, reversibly, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide?
- 12) A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound?
- 13) An unknown compound was found to have a percent composition as follows: 47.0% potassium, 14.5% carbon, and 38.5% oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula?
- 14) Rubbing alcohol was found to contain 60.0% carbon, 13.4% hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of the rubbing alcohol?