3. How many moles is 2.80×10^{24} atoms of silicon?

4. How many moles is 2.17×10^{23} representative particles of bromine?

5. How many atoms are in 1.14 mol of sulfur trioxide (SO_3)?

6. How many carbon atoms are in 2.12 mol of propane (C_3H_8)?

    How many hydrogen atoms are in 2.12 mol of propane(C_3H_8)?

13. How many moles is 1.50×10^{23} molecules of NH_3?

14. How many atoms are in 1.75 mol of $CHCl_3$?

15. What is the molar mass of $CaSO_4$?

51. Name the representative particle (atom, molecule, or formula unit) of each substance.

      a. oxygen gas

      b. sodium sulfide

      c. sulfur dioxide

      d. potassium

52. How many hydrogen atoms are in a representative particle of each substance?

      a. $Al(OH)_3$

      b. $H_2C_2O_4$

      c. $(NH_4)_2HPO_4$

      d. $C_4H_{10}O$

54. Find the number of moles in each substance.

      a. 2.41×10^{24} formula units of NaCl

      b. 9.03×10^{24} atoms of Hg

      c. 4.65×10^{24} molecules of NO_2

55. Which contains more molecules: 1.00 mol C_2H_2, 1.00 mol C_2H_6, or 1.00 mol CO?

56. Which contains more atoms: 1.00 mol C_2H_2, 1.00 mol C_2H_6, or 1.00 mol CO?

57. Find the number of representative particles in each substance.

      a. 3.00 mol Sn

b. 0.400 mol KCl

c. 7.50 mol SO_2

d. 4.80 X 10^{-3} mol NaI