

Unit 9 – Chapter 15: Strong Acid-Strong Base Titration

Name _____

Period _____

A 25.0-mL sample of 0.100 *M* HNO₃ is titrated with 0.100 *M* NaOH. Calculate the pH after the addition of the following volumes of NaOH. Plot the results of your calculations as pH versus milliliters of NaOH on the back of this sheet.

1) 0.0 mL of NaOH

7) 24.5 mL of NaOH

2) 4.0 mL of NaOH

8) 24.9 mL of NaOH

3) 8.0 mL of NaOH

9) 25.0 mL of NaOH

4) 12.5 mL of NaOH

10) 25.1 mL of NaOH

5) 20.0 mL of NaOH

11) 26.0 mL of NaOH

6) 24.0 mL of NaOH

12) 28.0 mL of NaOH

13) 30.0 mL of NaOH

