Unit 9 – Chapter 15: Weak Acid-Strong Base Titration   **Name _____**

Period _____

A 25.0-mL sample of 0.100 M propanoic acid ($HC_3H_5O_2$) is titrated with 0.100 M NaOH. Calculate the pH after the addition of the following volumes of NaOH. Plot the results of your calculations as pH versus milliliters of NaOH on the back of this sheet.

1) 0.0 mL of NaOH   7) 24.5 mL of NaOH

2) 4.0 mL of NaOH   8) 24.9 mL of NaOH

3) 8.0 mL of NaOH   9) 25.0 mL of NaOH

4) 12.5 mL of NaOH   10) 25.1 mL of NaOH

5) 20.0 mL of NaOH   11) 26.0 mL of NaOH

6) 24.0 mL of NaOH   12) 28.0 mL of NaOH

                                                                                           13) 30.0 mL of NaOH