

CHAPTER 9 NAMING COMPOUNDS PRACTICE

Name _____ Period _____ Date _____

Name the following compounds in column 1; in column 2, give the chemical formula.

- | | |
|---|-----------------------------|
| 1. CdSO_4 | 21. copper (I) oxide |
| 2. FeCl_3 | 22. aluminum carbonate |
| 3. CoCO_3 | 23. lead (IV) acetate |
| 4. $\text{Ba}(\text{NO}_2)_2$ | 24. iron (II) hydroxide |
| 5. $\text{Zn}(\text{ClO}_3)_2$ | 25. zinc bromide |
| 6. $(\text{NH}_4)_2\text{S}$ | 26. magnesium phosphate |
| 7. AgBr | 27. iron (III) sulfate |
| 8. SnO_2 | 28. bismuth nitrate |
| 9. $\text{Ca}_3(\text{PO}_4)_2$ | 29. potassium nitrite |
| 10. $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$ | 30. copper (I) chlorate |
| 11. $\text{Fe}(\text{NO}_2)_2$ | 31. calcium chlorite |
| 12. Cu_2SO_3 | 32. magnesium nitride |
| 13. MnI_2 | 33. silver sulfide |
| 14. Bi_2S_3 | 34. sodium sulfite |
| 15. KClO_2 | 35. carbon monoxide |
| 16. Na_3PO_3 | 36. hydrofluoric acid |
| 17. CaF_2 | 37. nitric acid |
| 18. $\text{Sb}(\text{OH})_3$ | 38. dinitrogen hexachloride |
| 19. K_2CrO_4 | 39. sulfurous acid |
| 20. NH_3 | 40. dihydrogen monoxide |