

**Solubility Rules Worksheet**

1. Classify each of the substances as being soluble or insoluble in water.

- a. KBr =
- b.  $\text{PbCO}_3$  =
- c. zinc hydroxide =
- d. sodium acetate =
- e. silver iodide =
- f. zinc carbonate =

- g. silver acetate =
- h. copper (II) sulfide =
- i.  $\text{Mg}_3(\text{PO}_4)_2$  =
- j. KOH =
- k.  $\text{NH}_4\text{OH}$  =
- l.  $\text{Hg}_2\text{SO}_4$  =
- m.  $\text{PbI}_2$  =

2. Identify the two new compounds which form if the solutions, as suggested by the following table, were mixed. CIRCLE the names of the compounds which would precipitate from the solutions.

	KBr	$\text{Na}_2\text{CO}_3$	CaS	$\text{NH}_4\text{OH}$
$\text{AgNO}_3$				
$\text{BaCl}_2$				
$\text{Al}(\text{NO}_3)_3$				
$\text{CuSO}_4$				