

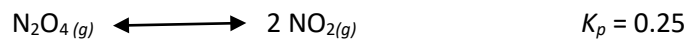
Unit 7 – Chapter 13: Chemical Equilibrium

Name _____

Assignment #3: Equilibrium & Partial Pressures with ICE Charts

Period _____

- 1) For the reaction



Calculate the equilibrium partial pressures of both species when the initial pressure of NO_2 is 0.050 atm.

- 2) At 35°C , $K_c = 1.6 \times 10^{-5}$ mol/liter for the reaction:



Calculate the concentrations of all species at equilibrium for each of the following original mixtures.

- 2.0 mol pure NOCl in a 2.0-liter flask
- 2.0 mol NO and 1.0 mol Cl_2 in a 1.0-liter flask
- 1.0 mol NOCl and 1.0 mol NO in a 1.0-liter flask
- 3.0 mol NO and 1.0 mol Cl_2 in a 1.0-liter flask
- 2.0 mol NOCl, 2.0 mol NO, and 1.0 mol Cl_2 in a 1.0-liter flask
- 1.0 mol/liter concentration of all three gases