

Unit 9 – Chapter 15: Acid-Base Equilibrium & Buffers

Name _____

Assignment #7: Solubility Calculations

Period _____

82) Calculate the solubility of each of the following compounds in moles per liter. Ignore any acid-base properties.



90) The K_{sp} for silver sulfate, Ag_2SO_4 , is 1.2×10^{-5} . Calculate the solubility of silver sulfate in each of the following:

a. water

b. 0.10 M AgNO_3

c. $0.20 \text{ M K}_2\text{SO}_4$

97) Will a precipitate form when 75.0 mL of 0.020 M BaCl_2 and 125 mL of $0.040 \text{ M Na}_2\text{SO}_4$ are mixed together?

98) Will a precipitate form when 100.0 mL of $4.0 \times 10^{-4} \text{ M Mg}(\text{NO}_3)_2$ is added to 100.0 mL of $2.0 \times 10^{-4} \text{ M NaOH}$?