

LAB: MAKING SALT

NAME _____

PERIOD _____

PURPOSE: To make a predetermined amount of sodium chloride from the proper amount of sodium hydrogen carbonate and hydrochloric acid.

SAFETY: Goggles and aprons must be worn.

COMPUTATION:

To make _____g (your assigned mass) of sodium chloride, determine the amount of sodium hydrogen carbonate needed by solving a mass-mass problem using the following balanced equation. Show your work.



PROCEDURE:

Weigh an evaporating dish. Weigh the calculated amount of sodium hydrogen carbonate in the evaporating dish, add hydrochloric acid until the bubbling stops and add a little more to insure a complete reaction. If time allows, set up a double boiler to evaporate to dryness or set aside to dry. Weigh the dry sodium chloride and evaporating dish. Find the experimental error. Find the percent of error, check your results on the computer and discuss the possible sources of error.

DATA:

1. Mass of evaporating dish + sodium hydrogen carbonate _____g
2. Mass of evaporating dish _____g
3. Mass of sodium hydrogen carbonate _____g
4. Mass of evaporating dish + sodium chloride _____g
5. Mass of evaporating dish _____g
6. Mass of sodium chloride (actual or experimental) _____g
7. Mass of sodium chloride (assigned or theoretical) _____g
8. Experimental error _____g
9. Percent of error (Show your steps) _____%
10. I have checked my results on the computer _____yes, _____no
11. Sources of error.