

15 PTS
(1/2 PT EACH)

CHAPTER 9 NAMING COMPOUNDS PRACTICE

Name KEY

Period _____ Date _____

Name the following compounds in column 1; in column 2, give the chemical formula.

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|--|---|
| 1. CdSO_4 CADMIUM SULFATE | 21. copper (I) oxide Cu_2O |
| 2. FeCl_3 IRON (III) CHLORIDE | 22. aluminum carbonate $\text{Al}_2(\text{CO}_3)_3$ |
| 3. CoCO_3 COBALT (II) CARBONATE | 23. lead (IV) acetate $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$ |
| 4. $\text{Ba}(\text{NO}_2)_2$ BARIUM NITRITE | 24. iron (II) hydroxide $\text{Fe}(\text{OH})_2$ |
| 5. $\text{Zn}(\text{ClO}_3)_2$ ZINC CHLORATE | 25. zinc bromide ZnBr_2 |
| 6. $(\text{NH}_4)_2\text{S}$ AMMONIUM SULFIDE | 26. magnesium phosphate $\text{Mg}_3(\text{PO}_4)_2$ |
| 7. AgBr SILVER BROMIDE | 27. iron (III) sulfate $\text{Fe}_2(\text{SO}_4)_3$ |
| 8. SnO_2 TIN (IV) OXIDE | 28. bismuth nitrate $\text{Bi}(\text{NO}_3)_3$ |
| 9. $\text{Ca}_3(\text{PO}_4)_2$ CALCIUM PHOSPHATE | 29. potassium nitrite KNO_2 |
| 10. $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$ ALUMINUM ACETATE | 30. copper (I) chlorate CuClO_3 |
| 11. $\text{Fe}(\text{NO}_2)_2$ IRON (II) NITRITE | 31. calcium chlorite $\text{Ca}(\text{ClO}_2)_2$ |
| 12. Cu_2SO_3 COPPER (I) SULFITE | 32. magnesium nitride Mg_3N_2 |
| 13. MnI_2 MANGANESE IODIDE | 33. silver sulfide Ag_2S |
| 14. Bi_2S_3 BISMUTH SULFIDE | 34. sodium sulfite Na_2SO_3 |
| 15. KClO_2 POTASSIUM CHLORITE | 35. carbon monoxide CO |
| 16. Na_3PO_3 SODIUM PHOSPHATE | 36. hydrofluoric acid HF |
| 17. CaF_2 CALCIUM FLUORIDE | 37. nitric acid HNO_3 |
| 18. $\text{Sb}(\text{OH})_3$ ANTIMONY HYDROXIDE | 38. dinitrogen hexachloride N_2Cl_6 |
| 19. K_2CrO_4 POTASSIUM CHROMATE | 39. sulfurous acid H_2SO_3 |
| 20. NH_3 NITROGEN TRIHYDRIDE
(AMMONIA!) | 40. dihydrogen monoxide H_2O |