

Unit 9 – Chapter 15: Acid-Base Equilibrium & Buffers

Name _____

Assignment #6: Weak Acid-Strong Base Titration

Period _____

A 25.0-mL sample of 0.100 *M* propanoic acid ($\text{HC}_3\text{H}_5\text{O}_2$) is titrated with 0.100 *M* NaOH. Calculate the pH after the addition of the following volumes of NaOH. Plot the results of your calculations as pH versus milliliters of NaOH on the back of this sheet.

1) 0.0 mL of NaOH

7) 24.5 mL of NaOH

2) 4.0 mL of NaOH

8) 24.9 mL of NaOH

3) 8.0 mL of NaOH

9) 25.0 mL of NaOH

4) 12.5 mL of NaOH

10) 25.1 mL of NaOH

5) 20.0 mL of NaOH

11) 26.0 mL of NaOH

6) 24.0 mL of NaOH

12) 28.0 mL of NaOH

13) 30.0 mL of NaOH

